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Seasonal and diurnal characteristics of water soluble inorganic compounds in the gas and aerosol phase in the Zurich area

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Gas and aerosol samples were taken using a wet effluent diffusion denuder/aerosol collector (WEDD/AC) coupled to ion chromatography (IC) in the city of Zurich, Switzerland from August to September 2002 and in March 2003. Major water soluble inorganic ions; nitrate, sulfate, and nitrite were analyzed online with a time resolution of two hours for the gas and aerosol phase. The fraction of water soluble inorganic anions in PM₁₀ varied from 15% in August to about 38% in March. Seasonal and diurnal variations of nitrate in the gas and aerosol phase were observed with more than 50% of the total nitrate in the gas phase during August and more than 80% of nitrate in the aerosol phase during March exceeding the concentration of sulfate by a factor of 2. Aerosol sulfate, on the other hand, did not show significant variability with season. However, in the gas phase, the SO₂ concentration was 6.5 times higher in winter than in summer. Nitrous acid (HONO) also showed a diurnal variation in both the gas and aerosol phase with the lowest concentration (0.2–0.6 µg/m³) in the afternoon. The primary pollutants, NO, CO and SO₂ mixing ratios were often at their highest between 04:00–10:00 local time due to the build up of fresh vehicle emission under a nocturnal inversion.

1. Introduction

Air pollutants can cause adverse health effects (Gilli et al., 1988; Maynard, 2004; Pyne, 2002), loss of visibility through light scattering (Lee and Sequeira, 2002) and have the potential to cause other environmental damages (Morales et al., 1998). The effects of atmospheric aerosols depend on their size and chemical composition (Seinfeld and Pandis, 1998). Water soluble components comprise most of the aerosol fraction (Harrison et al., 2004; Querol et al., 2004). The two major inorganic water soluble compounds which play a key role in aerosol properties are sulfate and nitrate.

Non-sea salt sulfate is often the most dominant secondary inorganic aerosol component in the atmosphere. Sulfate particles are formed by oxidation of sulfur dioxide

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(SO₂) either in the gas phase or in water droplets (Finlayson-Pitts and Pitts, 2000). Most of the SO₂ emissions globally result from fossil fuel burning. For this reason, most of the aerosol produced from the oxidation of SO₂ is considered to be anthropogenic (Seinfeld and Pandis, 1998). In an increasing number of cases, nitrate is the most abundant component in fine particles (Magliano et al., 1999; Schaap et al., 2004). Nitrate may exist in the accumulation mode (mainly as ammonium nitrate), but may also exist in coarse particles due to deposition of nitric acid (HNO₃) on alkaline mineral or salt particles (Bassett and Seinfeld, 1984; Murphy and Thomson, 1997; Gard et al., 1998; Henning et al., 2003).

Sulfur and nitrogen containing aerosols are of interest to atmospheric studies due to their influence on radiative transfer and cloud formation, and hence on global climate (Charlson et al., 1992). Several authors reported the radiative and cloud formation effect of sulfate aerosol. However, until recently nitrate has not been considered in assessments of the radiative effects of aerosols. Van Dorland et al. (1997) suggested that radiative forcing due to ammonium nitrate is about one tenth of the sulfate forcing. However, the importance of aerosol nitrate could increase substantially over the next century. Observations and model results show that in regions of elevated NO_x and NH₃ emissions, such as Europe, and parts of North America, NH₄NO₃ aerosol concentrations may be quite high and actually exceed those of sulfate (Schaap et al., 2002; Christoforou et al., 2000; Wieprecht et al., 2004).

The composition of PM₁₀ across Europe has been subject of a number of studies (Putaud et al., 2004; Turnbull et al., 2000; Querol et al., 2001; Pakkanen et al., 2001). In Switzerland, water soluble particulate inorganic compounds represent more than one-quarter of the average annual fine mass at both urban and rural sites (Hueglin et al., 2005). At urban sites in Switzerland, nitrate and sulfate account for ~14% and 15%, respectively, of the annual-average fine particle mass (PM₁₀). However, there is still a gap in measurements of gas phase precursors of these aerosol particulates, particularly nitric acid.

This paper presents results from two field campaigns conducted in the center of

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Zurich in summer 2002 and winter 2003. The study involves on-line quasi-continuous chemical characterization of particulate and gas-phase inorganic anions using a wet effluent diffusion denuder/aerosol collector (WEDD/AC) coupled to an ion chromatograph. This method is particularly of interest to nitrate measurements as it minimizes the artifact compared to filter sampling which often has artifacts due to vaporization (Schaap et al., 2004). In addition, the high time resolution of the online measurement helps to understand the diurnal variations of the gas and aerosol phase of the compounds which can be used for modeling.

2. Experimental

2.1. Sampling site

The city of Zurich is located in the north-eastern part of Switzerland. In 2003, Zurich city had a population of about 340 000 and covered a total area of about 88 km². The sampling was made at the Kaserne, an open space of ~1000 m² which is situated in the city center very close to the main train station (Zurich HB). The site is also used as air monitoring station by the Swiss national air pollution monitoring network (NABEL) and characterized as urban background site (Putaud et al., 2004; Hueglin et al., 2005). The sampling for this study was performed in a measuring van, located close to the NABEL station. Both the measuring van as well as the NABEL station were operated by the Swiss Federal Laboratories for Materials Testing and Research (EMPA). Two sampling campaigns were made, the first one from 8 August to 30 September, 2002 and the second from 3 to 30 March, 2003.

2.2. Methods

Sampling for chemical analysis was performed with a wet effluent diffusion denuder/aerosol collector (WEDD/AC) which contains a flattened glass tube denuder with a dimension of 350 mm × 3 mm × 30 mm (L × H × W) and an aerosol mixing chamber

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as described in Simon and Dasgupta (1993) and Zellweger et al. (1999). The inner surface of the denuder was coated with sodium silicate ($\text{Na}_2\text{Si}_4\text{O}_3$) and continuously wetted with ultra pure water ($18\text{ M}\Omega\text{-cm}$, Millipore Milli-Q) at a flow rate of 1.5 ml/min using an Ismatec peristaltic pump in counter flow to the air. At the end of the denuder, the effluent was pumped to a trace concentrator column.

The particle collection system is based on the growth of particles under supersaturated conditions. In a mixing chamber (volume 250 ml) heated water vapor (100°C , flow rate 0.6 ml/min) was mixed with the air stream. The air is pumped further to a maze impactor, which is cooled down to 2°C . The water vapor condenses on aerosol particles and grows to droplets, which impact on the maze. The particle containing liquid was collected and concentrated on a trace concentrator column before analysis by ion chromatography (IC). A more detailed description is found in Simon and Dasgupta (1995). Air samples were taken at a flow rate of 4 l/min through the WEDD/AC for 1 h , and the resulting effluents from both the denuder and aerosol collector were analyzed alternately using a Dionex ion chromatography system summarized in Table 1. Alternating analysis of the denuder and the particle collection system results in an overall temporal resolution of 2 h .

For the total carbon content, aerosol samples were collected on quartz fibre filters (150 mm diameter, Schleicher & Schuell) with a high-volume sampler and a PM10 inlet at a flow rate of $0.5\text{ m}^3/\text{min}$. Filters were sub-sampled by cutting out circles with a diameter of 14 mm . An automated C/N analyzer (EA-1108, Carlo Erba) was used for the determination of total carbon (TC).

Aerosol size distributions were determined with a scanning mobility particle sizer (SMPS) (TSI Model 3071A) from 3 to $30\text{ }\mu\text{m}$ of March. The SMPS determined the size and number of the aerosols for mobility diameters between 14 and 740 nm . The time resolution of the SMPS analysis was approximately 5 min , from which 1-h averages were calculated. In addition, meteorological parameters, ozone, NO , NO_2 , CO , SO_2 and PM10 were routinely determined at this station throughout the year. Data were provided by EMPA. Table 2 shows the methods and instruments used for the measure-

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ments of the above mentioned gaseous species and PM10.

2.3. Experimental uncertainties

The limits of detection for all the inorganic anions determined by IC were in the range 0.2–3.5 ng, which corresponds to air concentrations of about 1–20 ng/m³ for a sampling time of 1 h. In general, the concentrations of the ions determined by IC were sufficiently above the limit of detection. In some cases, when the concentrations of HNO₃ was small (<0.2 µg/m³) the co-elution of the nitrate peak with carbonate made the quantification difficult. In these cases, the concentration of the HNO₃ and aerosol nitrate could be overestimated by about 50%.

3. Results and discussion

SO₂ values from the WEDD/AC were compared to the NABEL data obtained by UV fluorescence at the same site (Fig. 1). In general the two methods showed good agreement ($r^2=0.71$), however, the UV fluorescence data seem to be overestimated by about 30%.

3.1. PM composition

A complete mass closure of PM10 in Switzerland is given by Hueglin et al. (2005). Figure 2 shows the chemical composition of PM10 measured at Zurich Kaserne in this study over the three months sampling period. In this plot TC (total carbon) refers to the sum of organic and elemental carbon. Sulfate was assumed to be present as ammonium sulfate, and particulate nitrate as ammonium nitrate. Chloride was assumed to be in the form of sodium chloride, and the mass of organic carbon was converted to a mass of organic compounds using a factor of 1.4 for urban background based on Harrison et al. (2003). It is expected that this conversion factor experiences a seasonal dependence caused by changes in the relative contributions of primary and secondary

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organic compounds. Due to the absence of data to support the use of a seasonally varying ratio, a constant factor throughout the year was assumed. The designation “unknown” applies to the difference between mass concentration determined by the beta gauge and the sum of the species concentrations identified by chemical analysis.

5 Silicates and calcium containing mineral dust, trace elements and water are among the materials not measured by the methods used in this study.

The composition of PM10 varied considerably from season to season. Carbonaceous aerosols represented the highest fraction (33% in August, 45% in September and 22% in March) except for March where the nitrate contribution was slightly higher. 10 The seasonal average fraction of water soluble inorganic compounds varied from about 15% in August to about 38% in March, with intermediate fractions in September. The percentage of nitrate in PM10 increased from summer to winter, showing a correlation with PM10 in March (Fig. 3). A similar observation was also reported by Querol et al. (2001). On the other hand, sulfate did not show a significant change in the contribution to the total PM10. The maximum contribution of sulfate was observed during 15 March (13%) and the minimum in August (10%), and no correlation with PM10 is seen (Fig. 3).

The current Swiss limit value for airborne particulate matter comprises both an annual mean concentration ($20\mu\text{g}/\text{m}^3$) and a daily concentration of $50\mu\text{g}/\text{m}^3$ which 20 should not be exceeded more than once per year. This daily concentration limit was not exceeded during August and September. However, in March the concentration of PM10 was above the limit for at least three days (Fig. 3). In March, PM10 accumulation was associated with a lower height of the mixing layer, higher emissions due to room heating, and a higher particulate nitrate formation due to lower temperature as 25 discussed below.

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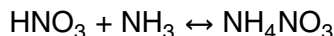
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3.2. HNO₃ and aerosol nitrate

3.2.1. Diurnal variations

The diurnal variations of particulate ammonium nitrate and gaseous HNO₃ are shown in Fig. 4. In August, the two compounds were found to be about equally abundant.

5 The average observed concentration of nitric acid for August was 0.9 μg/m³ reaching a maximum of 1.6 μg/m³ in the afternoon when solar radiation was strongest and temperatures were highest. The aerosol nitrate peaked slightly earlier. Both components show a decrease in the afternoon due to dry deposition of HNO₃ and an increase in early morning probably due to heterogeneous formation of nitrate. The maximum particulate nitrate concentration thus occurred in Zurich at 12:00 local time, which is four
10 hours later than in Pittsburgh (Wittig et al., 2004). This time lag is explained by the temperature dependent partitioning of ammonium nitrate between the gas and aerosol phase. To get a rough estimate of the ammonium nitrate partitioning during the maximum temperature, the expression given by Stelson and Seinfeld (1982) for the equilibrium constant (*K_p*) of the reaction between HNO₃ and NH₃ to form NH₄NO₃, was taken
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The equilibrium constant, *K_p*, of this reaction is given as the product of the partial pressures of NH₃ and HNO₃ by

$$20 \quad K_p = p_{\text{NH}_3} p_{\text{HNO}_3} . \quad (1)$$

If $K_p[\text{ppb}^2] < p_{\text{NH}_3} * p_{\text{HNO}_3}$, the system is supersaturated with ammonia and nitric acid, and a fraction of them will be transferred to the aerosol phase to establish equilibrium.

K_p is temperature dependent and can also be expressed as (Seinfeld and Pandis, 1998)

$$25 \quad \ln K_p = 84 - (24220/T) - 6.1 \ln(T/298) , \quad (2)$$

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where T is the temperature in Kelvin and for a total pressure of 1 atm.

K_p was calculated from Eq. (2) at $T=22^\circ\text{C}$ and fitted in Eq. (1) together with the measured HNO_3 concentration (0.6 ppb). Based on this equation, 11.8 ppb of ammonia is required to reach equilibrium. If the concentration of ammonia or nitric acid exceeds this value, the equilibrium shifts towards the right resulting in formation of ammonium nitrate. The calculated ammonia was compared with measurement values given by Emmenegger et al. (2004) for a tunnel entrance in Zurich city. The measured concentration at the tunnel entrance was much higher (36 ppb) compared to the calculated value, which implies that the formation of nitrate in Zurich city is very likely due to high ammonia concentrations.

3.2.2. Seasonal variations

A strong seasonal pattern of the NH_4NO_3 and nitric acid was observed. The highest contribution of NH_4NO_3 to total nitrate was found in March and the lowest in August. The contribution of nitrate to the PM_{10} was also 5 times higher in March than in August. The total nitrate concentration (sum of nitrate and nitric acid) was also substantially higher in winter than in summer. Since photochemistry is less efficient during winter, the increase in total nitrate concentrations in winter is explained by efficient heterogeneous formation of nitrate on deliquesced aerosol particles and a longer lifetime of nitrate in winter than in summer. The concentration of nitrate in winter also showed a good correlation with submicron particles ($\text{PM}_{0.74}$, see Fig. 5) ($r^2=0.52$) determined from the integrated SMPS spectra assuming spherical particles with a density of 1.5 g/cm^3 (ammonium nitrate and sulfate have densities of 1.725 and 1.77 g/cm^3 , respectively, which is reduced to some extent by carbonaceous compounds and water). This confirms that nitrate contributes substantially to $\text{PM}_{0.74}$ in winter (Fig. 5).

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3.2.3. Comparison with model results

The gas-particle partitioning of nitrate during March 2003 was further investigated using the Aerosol Inorganics Model II (AIM-II; <http://www.hpc1.uea.ac.uk/~e770/aim.html>; Clegg et al., 1998; Wexler and Clegg, 2002). The AIM-II describes the $\text{H}^+ - \text{NH}_4^+ - \text{SO}_4^{2-} - \text{NO}_3^- - \text{H}_2\text{O}$ system including solid phases and partitioning of NH_3 , HNO_3 and H_2SO_4 to the gas phase. The AIM determines the aerosol equilibrium state for a given amount of the above species by minimization of the Gibbs free energy of the system without simplifying assumptions. In this study sulfate, total nitrate, temperature and relative humidity measurements were available, and the aerosol was assumed to be fully neutralized with ammonia, that is $[\text{NH}_4^+] = [\text{NO}_3^-] + 2[\text{SO}_4^{2-}]$ in the model input. The aerosol was further assumed to be present as an aqueous solution, i.e. the formation of solids was turned off in the model calculations presented here. Model predictions including the solid phases were also made but predictions were similar to those for the aqueous aerosol.

The time series of the measured and modeled fraction of total nitrate in the particle phase along with the total nitrate, temperature and relative humidity is shown in Fig. 6. Low fractions of particulate nitrate were observed when the total nitrate loading was low, e.g. evening of 6 March, morning of 10 March, and 12 March, whereas high fractions of particulate nitrate coincided with high total nitrate loadings, e.g. from 19 March to 27 March. This reflects that NH_3 and HNO_3 start to partition to the particle phase in significant amounts only if their gas phase saturation concentration is exceeded. The dominating effect of varying total nitrate loadings is possibly one reason that no clear influence of the diurnal variations of temperature and relative humidity was observed, whereas diurnal cycles of ammonia concentrations might also be important. The model prediction captures the major trends of the nitrate partitioning reasonably well. However, the predicted particulate nitrate fraction is often lower than measured, and differences are typically larger during daytime. A main cause of that is probably that the atmospheric abundance of total ammonia is larger than the concentrations used in the

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model calculations thus driving more nitrate into the particle phase than predicted. Total ammonia concentrations used in the model runs ranged from 1 to $7\text{ }\mu\text{g}/\text{m}^3$, whereas ambient concentrations reported in Switzerland range from $3\text{--}7\text{ }\mu\text{g}/\text{m}^3$ at the rural site Payerne (EMEP station CH02) and up to about $25\text{ }\mu\text{g}/\text{m}^3$ (gas phase ammonia only) at a tunnel entrance in Zurich (Emmenegger et al., 2004). Furthermore, the tunnel entrance measurements showed a clear diurnal cycle with higher ammonia concentrations during daytime. This indicates that excess ammonia expected to be present in Zurich shifts the nitrate partitioning to the particle phase. Unfortunately, the AIM does not allow modelling with excess ammonia, that is $[\text{NH}_4^+] > [\text{NO}_3^-] + 2[\text{SO}_4^{2-}]$, which would provide an idea of the level of gas phase ammonia concentrations needed to explain the observed nitrate partitioning. It has also to be mentioned that part of the difference between observation and model calculation might arise from positive artifacts in particulate nitrate measurements due to organic nitrates or due to a disequilibrium state during the night caused by nitric acid production in aqueous particles through hydrolysis of N_2O_5 .

3.3. Nitrous acid (HONO)

HONO was measured in both campaigns in both the gas and the aerosol phase. It is well known that artifact HONO is formed from NO_2 if an alkaline solution is used as scrubber solution (Spinder et al., 2003), however, with MQ water as a scrubber solution, artifact HONO from NO_2 is negligible (Simon and Dasgupta, 1995; Acker et al., 2004). In the gas phase, a diurnal variation of HONO is observed with an increasing concentration at night (Fig. 7). During the day the concentration of HONO decreased significantly reaching a minimum between $0.2\text{--}0.6\text{ }\mu\text{g}/\text{m}^3$. These values were comparable to the one reported by Kleffmann and Gavriloaiei (2005). Formation of HONO during night time has been observed in urban areas (Acker et al., 2004) and is attributed to heterogeneous reactions of NO_2 on various surfaces. This heterogeneous formation and accumulation processes during the night lead to maximum values in the morning hours before photolysis causes a significant decrease. A similar

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trend was also observed for particulate nitrite, however with a less persistent increase during the night. The ratio of aerosol phase nitrite to total nitrite varied from 25% in August to 41% in September. Gaseous HONO concentrations were positively correlated with NO₂, with an average HONO/NO₂ ratio of 2×10^{-2} (Fig. 8). This value is significantly higher than the primary ratio from vehicles of 8.6×10^{-3} as reported by Kurtenbach et al. (2001) indicating that heterogeneous formation processes such as the ones reported by Gutzwiller et al. (2002) contribute substantially to the total HONO concentration in Zurich.

3.4. Sulfate and sulfur dioxide

Figure 9 shows the diurnal variations of the measured particulate SO₄⁻² concentrations, the gas phase SO₂ concentration, and the ratios between these two. The diurnal variation of SO₂ was most pronounced in March. This is due to increased early morning emissions into a planetary boundary layer with a low mixing height. Other combustion products such as CO and NO also show similar variations over the day (Fig. 10). The strong seasonal variation of SO₂ concentrations with highest values in March is explained by higher emissions due to room heating, lower mixing height of the planetary boundary layer, and less efficient oxidation to sulfate. Similar observations are also reported by Ruuskanen et al. (2003) and Altshuller (1984).

The average concentration of sulfate during the whole measurement period was $2.3 \mu\text{g}/\text{m}^3$. The highest concentration of sulfate was observed during March (Table 3), however, the sulfate contribution to PM₁₀ stayed virtually constant (see above). The sulfate values shown here are lower than those reported by Hueglin et al. (2005) for the same site. This is explained by a general decrease in the SO₂ emission over the last years as shown by the NABEL data from Zurich (Fig. 11). The ratio of SO₂ to SO₄ was higher in winter than in summer (Fig. 9), due to lower concentrations of reactive species (OH, HO₂ or RO₂, O₃) (Dudkin et al., 2002). The same reason also explains the diurnal variation rather showing a maximum during the day (Fig. 9).

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4. Conclusions

Water soluble inorganic anions comprise on average more than one quarter of the total PM₁₀. Nitrate contributes a larger fraction to PM₁₀ in winter than in other seasons due to lower temperature. The value of nitrate also exceeds the one of sulfate during winter. The variations of nitrate were strongest, which is attributed to the temperature dependency of nitrate aerosol partitioning. In August with high ambient temperatures the nitric acid concentration was highest, and due to a higher deposition rate of HNO₃ than particulate nitrate, the total nitrate concentration in August was lower compared to March. The data showed marked seasonal and diurnal variations for gaseous HNO₃, HONO, SO₂, and their corresponding particulate phase except for sulfate.

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Table 1. Summary of the Dionex ion chromatography system used for online measurement of the anions.

Analytical column	AS 17, 4mm
Guard column	AG 17, 4mm
Pump	IP 20
Detector	CD 25 (conductivity detector)
Eluent generator	EG40 eluent generator
Concentrator column	TAC-LP1
Trap column	ATC-1
software	Peaknet 6.4
Suppressor	ASRS, 4mm
valve	Rheodyne

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Table 2. Instruments used by the NABEL for gas and PM10 measurements.

Compound	Method	Instrument
NO/NO ₂	Chemiluminescence	Horiba APNA 360
Ozone	UV-Absorption	TEI 49C
CO	Infrared-Absorption	Horiba APMA 360
PM10	Beta gauge	Digitel DA-80H
SO ₂	UV fluorescence	Monitor Labs 9850

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Table 3. Mean, median and standard deviations of the measured gas and aerosol compounds and PM10. The concentrations are all in $\mu\text{g}/\text{m}^3$ except CO which is in mg/m^3 . G = gas phase and A = aerosol phase.

August									
	G-NO3	G-SO2	A-NO3	A-SO4	NO2	NO	O3	CO	PM10
Mean	0.81	0.68	0.83	1.71	28.6	7.4	49.1	0.37	22.9
Median	0.55	0.41	0.40	1.53	25.3	3.0	46.3	0.33	22.6
StDv	0.85	0.80	1.10	1.32	15.3	11.9	34.1	0.17	9.3
September									
	G-NO3	G-SO2	A-NO3	A-SO4	NO2	NO	O3	CO	PM10
Mean	0.76	1.22	2.63	1.65	30.1	8.7	37.0	0.39	21.6
Median	0.50	0.90	1.39	1.25	28.2	4.3	30.3	0.35	20.5
StDv	0.71	1.18	3.22	1.43	13.9	15.1	26.9	0.16	10.5
March									
	G-NO3	G-SO2	A-NO3	A-SO4	NO2	NO	O3	CO	PM10
Mean	0.69	4.07	6.85	3.37	50.6	16.5	41.8	0.51	34.5
Median	0.63	2.87	4.10	2.89	44.3	5.0	40.9	0.43	30.6
StDv	0.36	3.64	5.04	1.87	27.0	27.1	32.7	0.23	19.0

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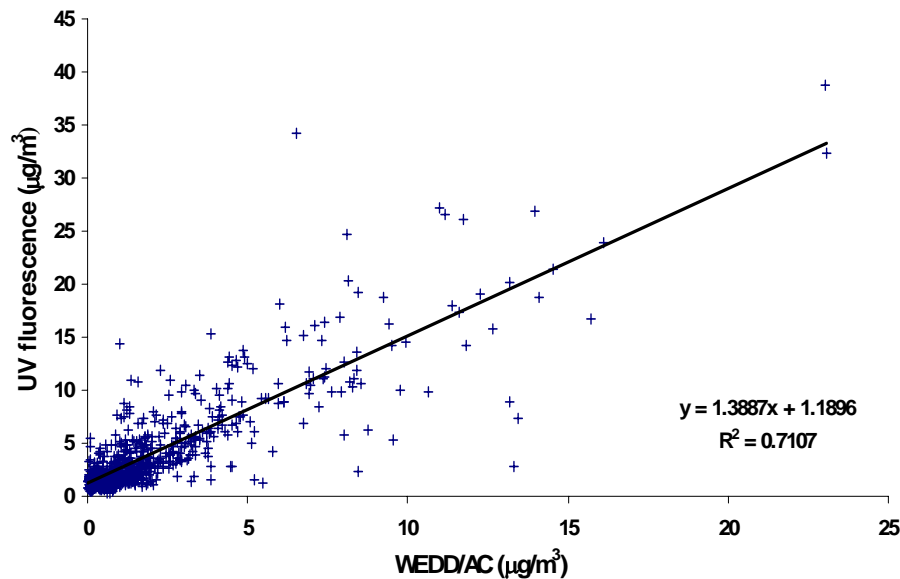


Fig. 1. Correlation between the WEDD/AC and UV-fluorescence method for SO_2 measurement.

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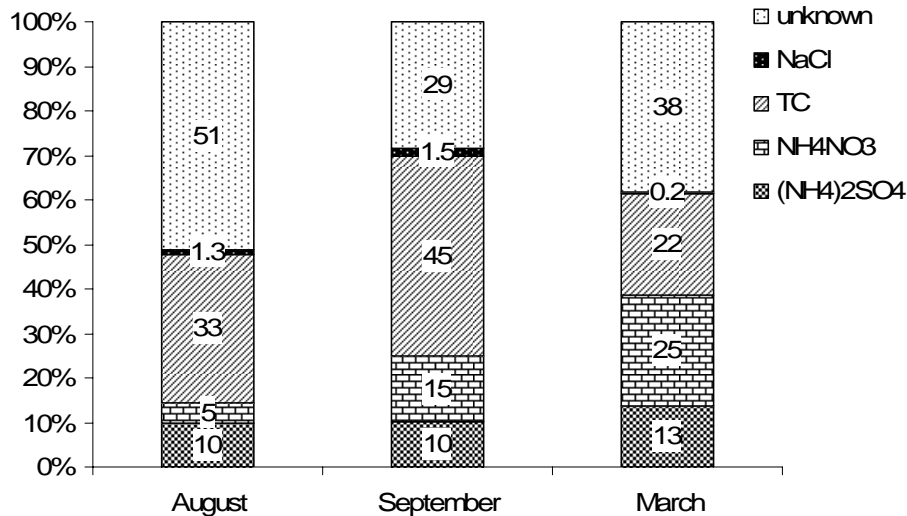


Fig. 2. Chemical composition of PM10 at Zurich Kaserne for the three months. The values in the bars indicate the percentage contributed by the component.

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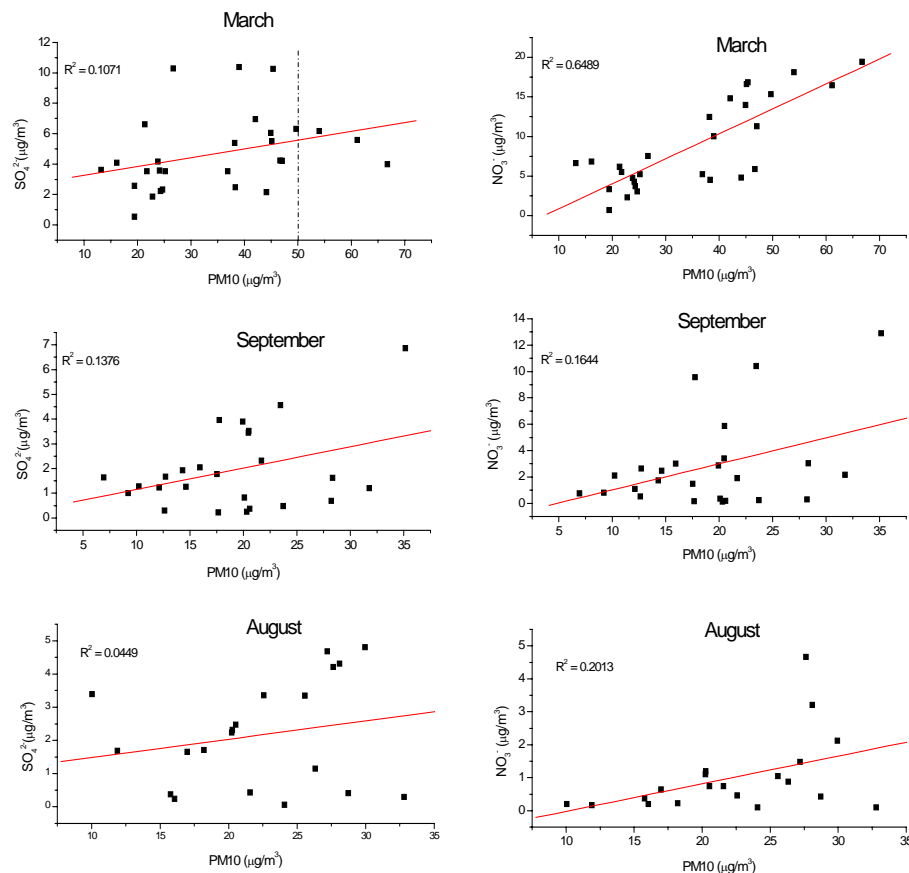


Fig. 3. Correlations of sulfate and nitrate (calculated as ammonium sulfate and ammonium nitrate, respectively) with PM_{10} . The dotted line in the March sulfate graph indicates the current Switzerland daily concentration limit value for airborne particulate matter, which is $50 \mu\text{g}/\text{m}^3$.

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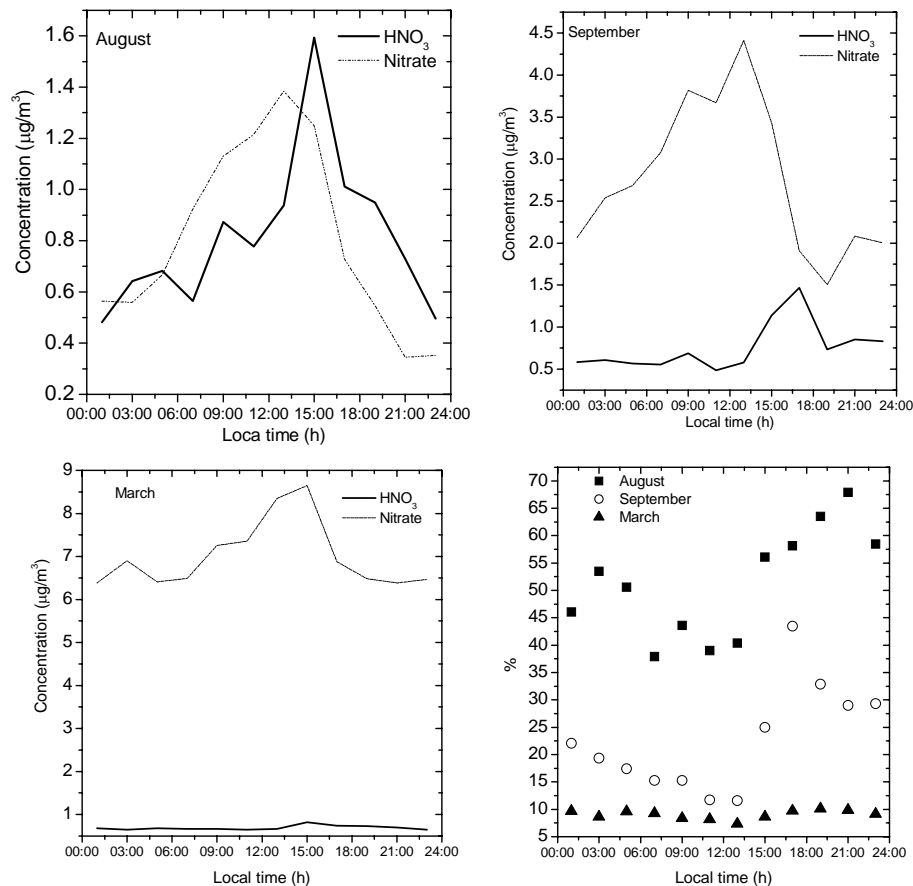


Fig. 4. Average diurnal variation of nitric acid and aerosol nitrate for the three months. The bottom right panel shows the contribution of HNO_3 of gaseous nitric acid to total nitrate in percent.

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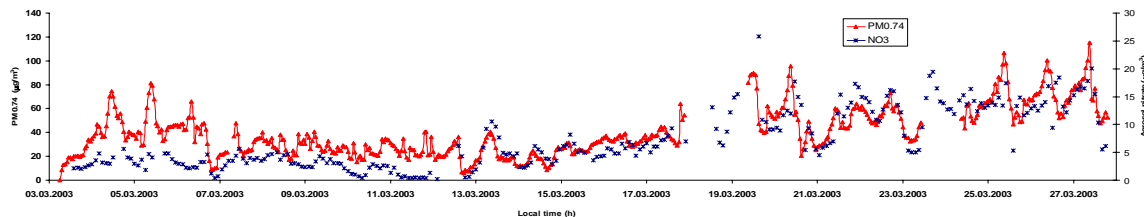


Fig. 5. Time series of PM0.74 and aerosol nitrate for March 2003.

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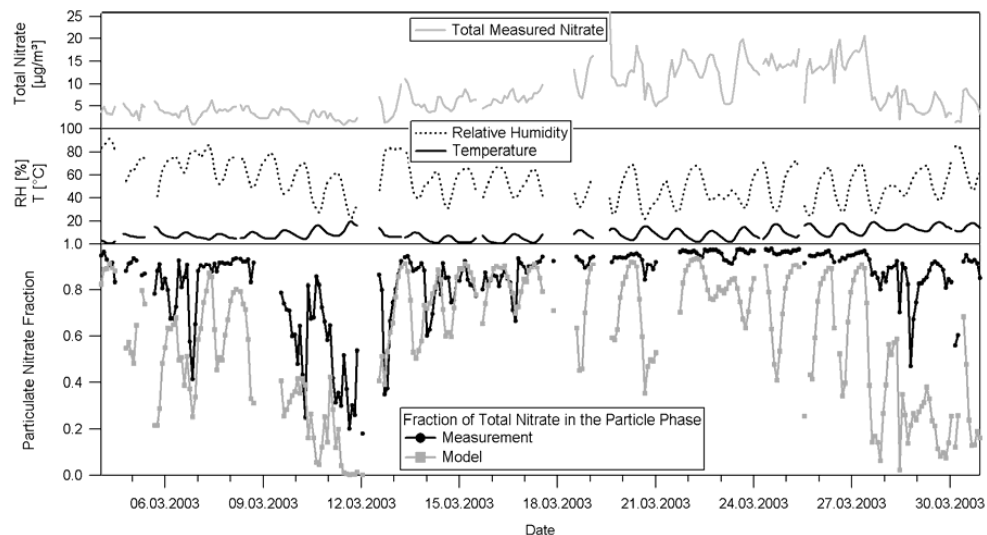


Fig. 6. Measured and modeled partitioning of nitrate between particle and gas phase (bottom panel). Total measured nitrate (top panel) and temperature and relative humidity (center panel) are also shown.

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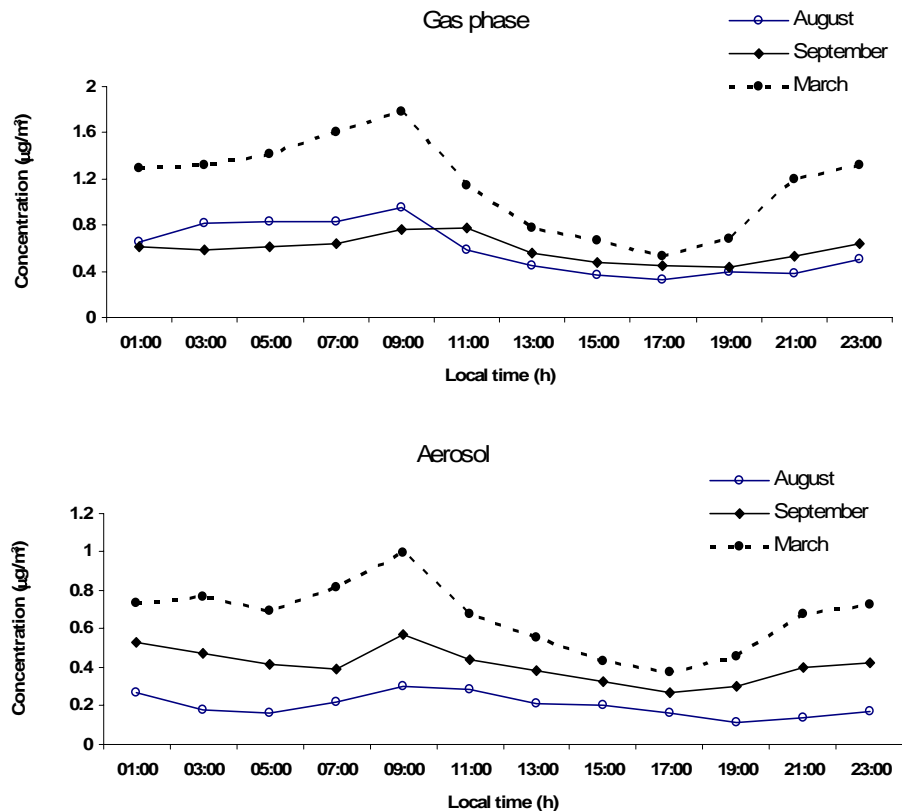
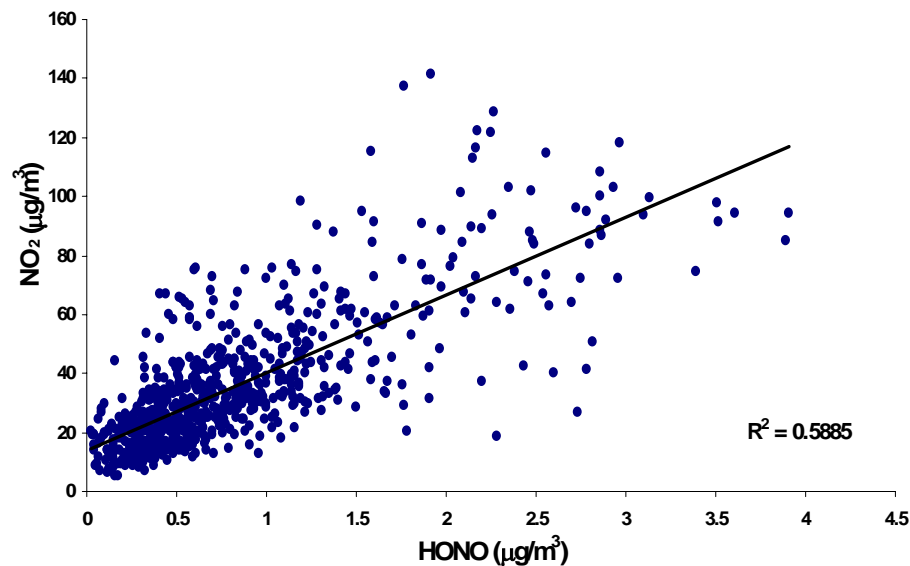


Fig. 7. Average diurnal variations of HONO in the gas and aerosol phase for the three months.

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**Fig. 8.** Correlation of HONO with NO₂.

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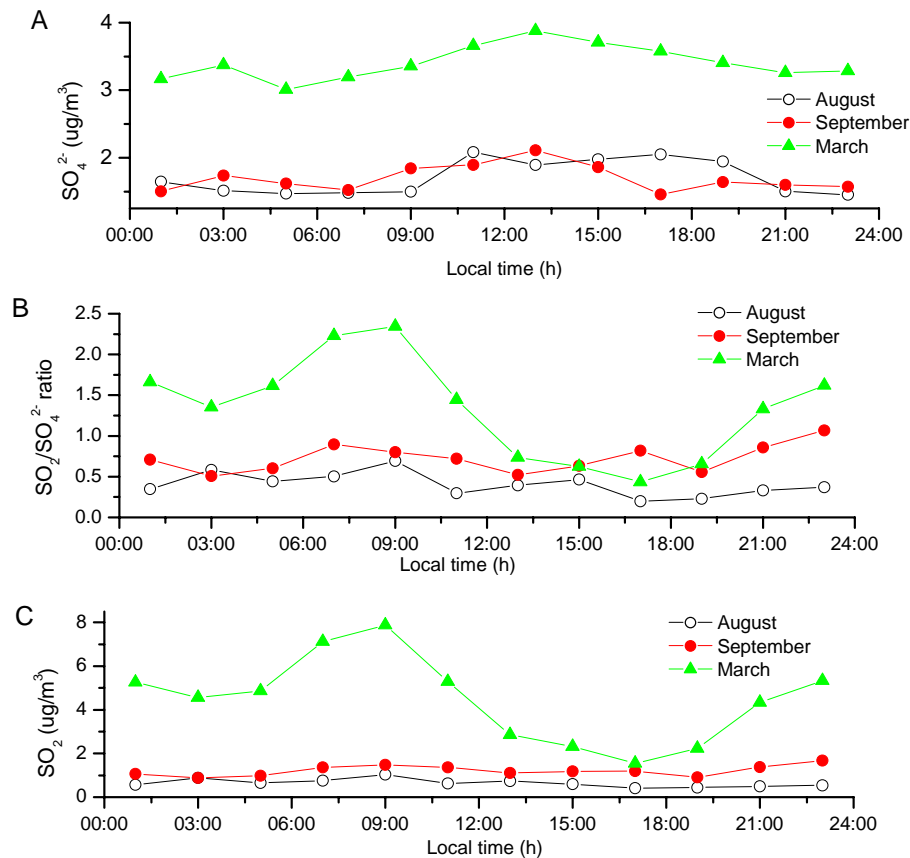


Fig. 9. (A) Diurnal variation of the average concentration of SO₄²⁻ for the three months. (B) Diurnal variation of SO₂/SO₄²⁻. (C) Diurnal variation of the average concentration of SO₂.

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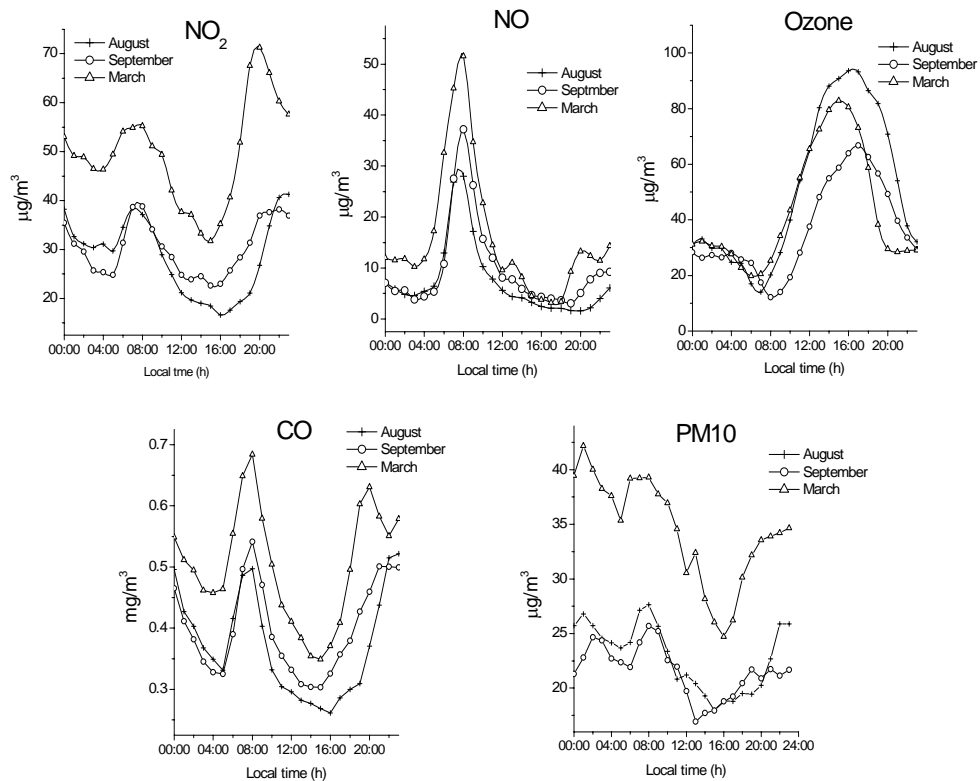
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**Fig. 10.** Diurnal variations of gaseous pollutants and of PM₁₀ for the three months.

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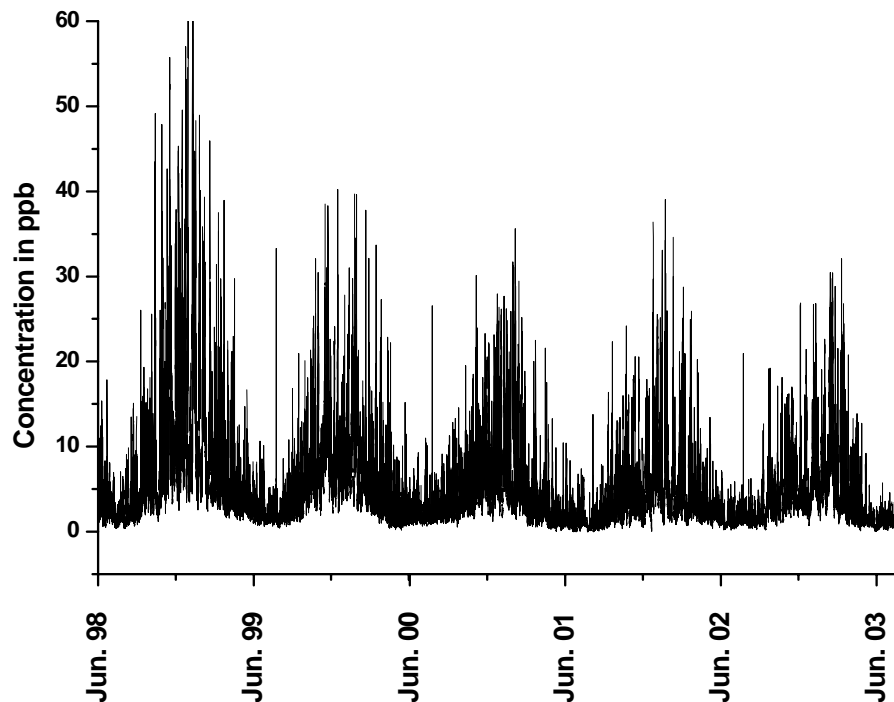


Fig. 11. Concentration of SO₂ over a five years' period.

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